

Acids, Bases, and Solutions ▪ *Review and Reinforce***Concentration and Solubility****Understanding Main Ideas**

Answer the following questions on the lines below.

1. What amounts do you compare when measuring concentration?

2. How can you tell that a white powder is salt without tasting it?

3. Which solution will have more gas dissolved in it, a solution under high pressure or one under low pressure?

4. Explain the meaning of the expression “like dissolves like.”

5. How does temperature affect the solubility of most solids?

Building Vocabulary

Match each term with its definition by writing the letter of the correct definition on the line beside the term in the left column.

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| _____ 6. dilute solution | a. a measure of how much solute can dissolve in a solvent at a given temperature |
| _____ 7. concentrated solution | b. a solution that has more dissolved solute than is predicted by its solubility |
| _____ 8. solubility | c. a solution that has so much solute that no more dissolves |
| _____ 9. saturated solution | d. a solution that has only a little solute |
| _____ 10. unsaturated solution | e. a solution in which more solute can be dissolved |
| _____ 11. supersaturated solution | f. a solution that has a lot of solute |