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| Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | Class: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date \_\_\_\_\_\_\_\_\_\_\_\_\_ |

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| https://ci4.googleusercontent.com/proxy/w1kKjQcK7Hqrc9l7FTSPWunyfotKPHvTYCPkk_gAitqwcWZoyEDImDOuzh1p9IWaesenEAHxu2tNWGLrro-9wyF0PQYr_NOjb1rYUS0uVRADDbWMnelgrpKJ6gOFKEM=s0-d-e1-ft#http://www.teach-nology.com/worksheets/soc_studies/amrev/read/reading.gif | **Reading Comprehension Worksheet**  **Compound & Elements** |

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| The 100 plus elements are the "alphabet" of matter because every tangible material is a combination of these elements.  The elements are categorized as: Metals, Non-Metals, Metalloids, Noble Gases  This main categorization is based on the electrical attributes of the elements. Some elements tend to "lose" an electron, and become positively charged. (These are the metals), Other elements tend to "acquire" an electron, and become negatively charged, (the nonmetals). The third group, (metalloids), falls in between these two extremes. And the Noble Gas category is "unelectrical" -- displays no interest in either acquiring and losing an electron, but unlike the metalloids who can be "persuaded" one way or the other -- the noble ones simply don't engage.  Since chemical reaction and chemical bonding are electrical in nature, it so happens that members of a certain category can substitute for another member and thereby create a combination which is slightly different, but generally similar. Say then that by carefully choosing a replacement element in a chemical compound, it is possible to "engineer" a desired slight change in the nature of the compound. This used to be the art, and now is the science of chemistry. |

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| 1. Which elements are likely to lose electrons? | \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |
| 2. Which elements are likely to gain electrons? | \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |
| 3. Which type of elements are likely to have no electrical charge at all? | \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |
| 4. How does knowing a great deal about the property of certain elements help us as humans? | \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |